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A substance combines with oxygen, releasing a large amount of energy as heat and light, in a(n) . 18. The decomposition of a substance by an electric current is called . 19. A(n) orders the elements by the ease with which they undergo certain chemical reactions. 20.

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Modern Chemistry 24 Chapter Test .

Name Class Date I Chapter Test B,

continued 7. The discovery of the nucleus was a result of Rutherford's observation that a small percentage of the positively charged particles bombarding the metal's surface a. were slightly deflected as they

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passed through the metal.

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Assessment Chapter Test B Teacher Notes
and Answers 2 Measurements and

Calculations TEST B 1. a 2. c 3. c 4. a 5. c

6. c 7. a 8. a 9. time 10. mass 11. density

12. energy or work 13. length or distance

14. volume 15. area 16. qualitative 17.

quantitative 18. qualitative 19. quantitative

20. 3.00 ± 10^5 km/s 21. three 22. 2.6 ± 10

2 2 g 23. 2.50 ± 10^2 L 24. quantity 25.

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Modern Chemistry 141 Chapter Test

Name Class Date Chapter Test B,

continued 18. When a strong acid is

titrated with a weak base, the pH of the

solution at the equivalence point is than 7.

19. A is a highly purified solid used to

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check the concentration of a standard solution. 20. A 1 M solution of NaOH will have a pH that is

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Holt McDougal Modern Chemistry
Chapter Test Assessment Chapter Test B
Teacher Notes and Answers 5 The
Periodic Law TEST B 1. a 2. c 3. d 4. d 5.
a 6. a 7. c 8. a 9. lanthanides 10. 2 11.
fourth 12. transition elements 13. 32 14.
valence electrons 15. electron affinity 16.
electronegativity 17. ionization energy 18.
3 s² 3p⁴ 19. atomic radius 20 ...

~~Assessment Chapter Test B~~ — Wag & Paws
Modern Chemistry 33 Chapter Test Name
Class Date Chapter Test B, continued 15.
The energy state of an atom is called its
ground state. 16. The number of waves
that pass a point in one second is called.
17. When an electron drops from a higher-

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energy state to a lower-energy state, a(n)
spectrum is produced. 18.

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Modern Chemistry 91 Chapter Test Name
Class Date Chapter Test B, continued 27.

Distinguish between ionic crystals and
metallic crystals. PART V Write the
answers to the following questions on the
line to the left, and show your work in the
space provided. Modern Chemistry
Chapter 12 Review Answer Key

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~~Name_Class_Date Assessment ...~~

Modern Chemistry 20 Chapter Test Name

Class Date Chapter Test A, continued

_____ 13. To calculate the number of atoms present in 2.0 mol of an element, you would a. add Avogadro's number of atoms per mole to 2.0 mole. b. subtract Avogadro's number of atoms per mole from 2.0 mole. c.

~~Modern Chemistry Chapter Atoms Test~~

~~Answers~~

Holt McDougal Modern Chemistry 3

Chapter Test Chapter Test B, continued 16

Modern chemistry chapter 3 test b

answers. The measure of the ability of an atom in a chemical compound to attract electrons from another atom in the

compound is called _____. 17. The energy required to remove one electron from an

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atom is called its macamp. 18.

~~Modern Chemistry Chapter 3 Test B~~ ~~Answers~~

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Aqueous Solutions and Colligative
Properties Chapter Practice Test Test your
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CHAPTER 3 TEST continued Date Class
FILL IN THE BLANK Write the correct
term (or terms) in the space provided. 9, If
a particular ompound is composed of
elements A and B, the ratio of the mass of
B to the mass of A will alw ys be the
same. This is a statement of the law of
exactly 12 g of carbon-12 is referred to as
a(n) 11.

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~~San Ramon Valley High School~~

(b) positive ions. (d) negative ions. 3. b Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is (a) higher in potential energy. (c) equal in potential energy. (b) lower in potential energy. (d) unstable. 4. b The lattice energy of compound A is greater in magnitude than that of ...

~~6 Chemical Bonding~~

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b. Explain how an atom can exist in this state. Atoms consist of a positively charged nucleus, made up of protons and

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neutrons, that is surrounded by a negatively charged electron cloud. The positive and negative charges combine to form a net neutral charge. MODERN CHEMISTRY ATOMS: THE BUILDING BLOCKS OF MATTER 19

~~3 Atoms: The Building Blocks of Matter~~

Modern Chemistry 137 Chapter Test
Name Class Date Chapter Test A,
continued ____ 19. In the figure on the
previous page, the pH at the equivalence
point a. is equal to 7.0. b. is greater than
7.0. c. is less than 7.0. d. cannot be
determined from the data given. ____ 20.
In the figure on the previous page, the
volume of titration standard

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CHAPTER 22 TEST Nuclear Chemistry
Class MULTIPLE CHOICE On the line at
the left of each statement, write the letter
of the choice that best completes the
statement or answers the question. After
converting units, the nuclear mass defect
is equivalent to the a. atomic mass b.
electrostatic force c. energy of chemical
reaction

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CHAPTER 16 REVIEW Reaction Energy
SECTION 2 SHORT ANSWER Answer
the following questions in the space
provided. 1. For the following examples,
state whether the change in entropy favors
the forward or reverse reaction: forward
reaction a. $\text{HCl}(l) \rightleftharpoons \text{HCl}(g)$ reverse
reaction b. $\text{C}_6\text{H}_{12}\text{O}_6(aq) \rightleftharpoons \text{C}_6\text{H}_{12}\text{O}_6(s)$ forward
reaction c. $2\text{NH}_3(g) \rightleftharpoons \dots$

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