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After Rutherford's scattering of α -particles experiment, following conclusions were made as regard as atomic structure : (a) Most of the mass and all of the charge of an atom concentrated in a very small region is called atomic nucleus. (b) Nucleus is positively charged and its size is of the order of 10^{-15} m

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The atomic structure is an important chapter, and it is likely to hold importance as it will prepare you for advanced concepts based on the atomic model. For instance, the bonding of elements with other elements via sharing electrons is included in organic chemistry, which is yet again an essential part of Chemistry as a whole.

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Chapter 2: Atomic Structure and Chemical Bonding. 1. Chapter 2 1. Chapter 2: Atomic Structure and Chemical Bonding. • Materials Molecules Atoms • Atoms = protons (p) + neutrons (n) + electrons (e) • Protons and neutrons are made of quarks • Quantitative measurements need units: metric or S.I.(Systeme International) or mks(meter-kilogram-second) units.

Chapter 2: Atomic Structure and Chemical Bonding

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The Structure of an Atom Protons are found in the nucleus of an atom, have a positive charge, and a mass of one atomic mass unit (AMU) each. The number of protons is equal to the atomic number on the periodic table and determines the identity of the element. Neutrons are also found in the nucleus.

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